

1 vii Chemistry
Multiple Choice Questions (MCQs)
(for 2nd Term)
CLASS: VII
SUBJECT: CHEMISTRY

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Chapter – 4

- Question 1) The word atom is derived from the Greek word –
(a) anu (b) paramanu (c) atomos (d) none of these
- Question 2) Matter is composed of very small particles called –
(a) atom (b) element (c) molecule (d) compound
- Question 3) Atoms of the same elements are –
(a) same (b) different (c) may be same or different (d) none of these
- Question 4) The fundamental particles of atom are –
(a) electron (b) proton (c) neutron (d) all of these
- Question 5) Where protons and neutrons are present in an atom?
(a) orbits (b) nucleus (c) in both a and b (d) none of these
- Question 6) Electrons are revolved around the nucleus in fixed path are called –
(a) orbits (b) nucleus (c) both a and b (d) none of these
- Question 7) The smallest particle of an element or a compound that has independent existence is called –
(a) atom (b) molecule (c) compound (d) none of these
- Question 8) Mass of an atom and its subatomic particles are measured in –
(a) kg (b) Jule (c) amu (d) none of these
- Question 9) Which charge is carried by neutron :
(a) positive (b) negative (c) both a and b (d) no change
- Question 10) Positively charged particle of an atom is –
(a) proton (b) electron (c) neutron (d) all of these
- Question 11) Negatively charged particle of an atom is –
(a) proton (b) electron (c) neutron (d) all of these
- Question 12) The combining capacity of an atom is –
(a) valency (b) atomicity (c) both a and b (d) none of these
- Question 13) The number of proton of an atom is called –
(a) atomic number (b) mass number (c) both a and b (d) none of these
- Question 14) Mass number is sum of –
(a) number of protons (b) number of electrons (c) number of neutron (d) both a and c
- Question 15) Electrons present in valence shell are called –
(a) electronegative (b) electropositive (c) valence electron (d) none of these
- Question 16) Which of the following is an electronegative monovalent radical?
(a) Sulphate (b) Sulphite (c) Carbonate (d) Bromine
- Question 17) Which of the following is an electronegative bivalent radical?
(a) Sulphate (b) Bromide (c) Chloride (d) iodide
- Question 18) Na⁺ is an example of –
(a) electronegative monovalent (b) electronegative bivalent
(c) electropositive monovalent (d) none of these
- Question 19) Example of electropositive divalent radical is –
(a) Calcium (b) Bromide (c) iodide (d) Chloride
- Question 20) Explain of electropositive trivalent radical is –
(a) Calcium (b) Bromide (c) Chromium (d) none of these
- Question 21) How many periods are there in periodic table?
(a) 3 (b) 4 (c) 7 (d) 8
- Question 22) How many vertical columns are there in periodic table?
(a) 17 (b) 16 (c) 7 (d) 18
- Question 23) Elements are arranged in periodic table according to increasing order of their –
(a) atomic number (b) mass number (c) chemical properties (d) none of these
- Question 24) Elements present in a group have same –
(a) valence electrons (b) properties (c) both a and b (d) none of these
- Question 25) What is the atomic number of magnesium element?
(a) 14 (b) 12 (c) 16 (d) 9

Chapter – 5

- Question 1) Two substances react together to form a new –
(a) molecule (b) compound (c) element (d) all of these
- Question 2) A process that involves rearrangement of the molecular or ionic structure of a substance is called –
(a) chemical reaction (b) chemical formulae (c) both a and b (d) none of these
- Question 3) Which of the following is characteristic of chemical reaction :
(a) change in colour (b) change of state (c) change in smell (d) all of these
- Question 4) During chemical reaction energy is
(a) absorbed (b) released (c) both a and b (d) none of these
- Question 5) What are the products of burning of candle?
(a) CO₂ (b) H₂O (c) both a and b (d) none of these
- Question 6) When sodium reacts with hydrochloric acid, it released –
(a) Hydrogen gas (b) carbon dioxide gas (c) both a and b (d) none of these
- Question 7) Potassium iodide reacts with lead acetate to form
(a) lead iodide (b) lead sulphate (c) lead nitrate (d) none of these

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- Question 8) During formation of slaked lime, heat is –
 (a) absorbed (b) released (c) both a and b (d) none of these
- Question 9) Which reaction takes place when two substances react together to form a new compound –
 (a) chemical reaction (b) physical reaction (c) reversible reaction (d) none of these
- Question 10) On heating solid ammonium chloride ammonia and hydrogen chloride gas are –
 (a) absorbed (b) released (c) both a and b (d) none of these
- Question 11) One molecule of calcium carbonate reacts with 2 molecules of hydrochloric acid to form –
 (a) one molecule of calcium chloride (b) one molecule of water
 (c) one molecule of carbon dioxide (d) all of these
- Question 12) Ammonia and hydrogen chloride gas are formed by heating –
 (a) ammonium chloride (b) sodium chloride (c) potassium chloride (d) none of these
- Question 13) Chemical involves the study of a large number of –
 (a) elements (b) compounds (c) mixtures (d) all of these
- Question 14) The process that leads to a chemical change is called a –
 (a) chemical reaction (b) chemical equation (c) physical reaction (d) none of these
- Question 15) The substances take part in a chemical reaction are called –
 (a) reactants (b) products (c) both a and b (d) none of these
- Question 16) The substances produced in a chemical reaction are called –
 (a) reactants (b) products (c) both a and b (d) none of these
- Question 17) The product of rusting of iron is –
 (a) milk (b) water (c) rust (d) none of these
- Question 18) When potassium permanganate reacts with citric acid, the colour of potassium permanganate –
 (a) changed (b) disappear (c) not changed (d) none of these
- Question 19) An arrow in a chemical reaction pointing towards –
 (a) reactants (b) products (c) catalyst (d) none of these
- Question 20) Iron reacts with oxygen to produce –
 (a) Iron oxide (b) iron sulphide (c) Iron sulphate (d) none of these
- Question 21) What is the formula of iron oxide ?
 (a) Fe_2O_3 (b) FeO_4 (c) Fe_3O_4 (d) Fe_2O_2
- Question 22) In charring of sugar, sugar is liberated into –
 (a) carbon (b) water (c) both a and b (d) none of these
- Question 23) A change cannot take place without the involvement of –
 (a) energy (b) reactant (c) product (d) all of these
- Question 24) Which of the following is balanced chemical equation?
 (a) $\text{KClO}_3 \rightarrow \text{KCl} + \text{O}_2$ (b) $\text{MnCl}_2 + \text{HCl} \rightarrow \text{MnCl}_2 + \text{H}_2\text{O}$
 (c) $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$ (d) none of these
- Question 25) Any chemical equation in which balanced elements does not exist is called a –
 (a) balanced equation (b) unbalanced equation (c) skeleton equation (d) both b and c

